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Key indicators

Single-crystal X-ray study
 $T = 293$ K
Mean $\sigma(\text{C}-\text{C}) = 0.003$ Å
 R factor = 0.040
 wR factor = 0.121
Data-to-parameter ratio = 17.9For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.2-(Benzotriazol-1-yl)-2-(α -hydroxybenzyl)-
 N -phenylthioacrylamide

In the title compound, $\text{C}_{21}\text{H}_{16}\text{N}_4\text{OS}$, the dihedral angles between the planes of the benzotriazole and N -phenyl rings and the plane of the atoms that link these two rings are $79.56(6)$ and $59.02(5)^\circ$, respectively, while that between the two benzene rings is $64.12(6)^\circ$. There are some inter- and intramolecular interactions in the crystal structure.

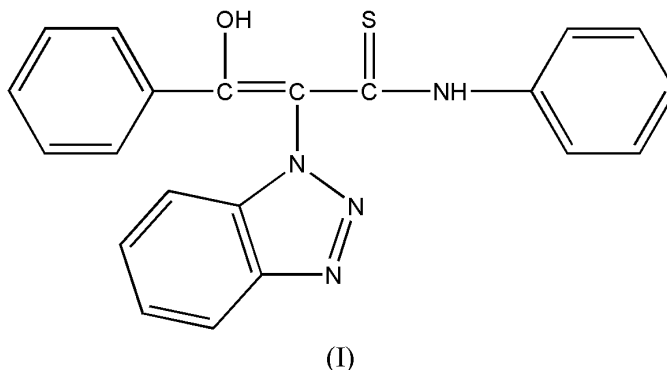
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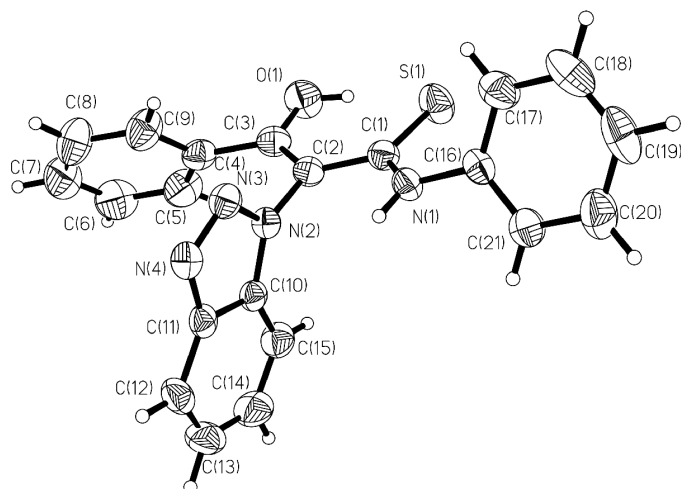
Comment

Triazole nuclei appear frequently in the structures of various natural products and biologically active compounds, notably thiamine (vitamin B), penicillins, and antibiotics such as micrococcin (James & Watson, 1966). Triazole derivatives have also attracted considerable attention in industry and agriculture, because of their significant biological activities (Zhang *et al.*, 2002). In this paper, we report the structure of the title compound, (I).



In compound (I) (Fig. 1), the bond lengths and angles are generally normal in the rings (Ji *et al.*, 2002). The $\text{C}=\text{S}$ bond length in (I) is close to the typical $\text{C}=\text{S}$ double-bond length (Table 1). Atom C2 lies in the plane of the benzotriazole ring, and atoms S1, N1, C1 and C2 are coplanar (plane $P1$). The dihedral angles formed by the benzotriazole and C16–C21 rings with $P1$ are $79.56(6)$ and $59.02(5)^\circ$, respectively. The $\text{C1}-\text{N1}-\text{C16}-\text{C17}$, $\text{N3}-\text{N2}-\text{C2}-\text{C1}$, $\text{N3}-\text{N2}-\text{C2}-\text{C3}$ and $\text{N2}-\text{C2}-\text{C3}-\text{C4}$ torsion angles are $54.8(2)$, $84.17(18)$, $92.39(17)$, $3.6(2)^\circ$, respectively.

The most interesting structural features of the title compound are the $\text{N}-\text{H}\cdots\text{N}$ inter- and intramolecular hydrogen bonds, and the strong $\text{O}-\text{H}\cdots\text{S}$ intramolecular interaction (Table 2). These interactions stabilize the structure of (I).

**Figure 1**

The structure of (I), with the atom-numbering scheme. Displacement ellipsoids are drawn at the 35% probability level and H atoms are shown as small spheres of arbitrary radii.

Experimental

A mixture of α -(benzotriazol-1-yl)acetophenone (1.34 g, 0.01 mol), phenyl isothiocyanate (2.42 g, 0.01 mol), potassium hydroxide (0.4 g, 0.01 mol) and dimethylsulfoxide (50 ml) was stirred for 1 h at room temperature. The solution was then filtered, concentrated and purified by flash chromatography (silica gel, chloroform:cyclohexane, 5:1) to afford the title compound (yield 3.01 g, 80%; m.p. 433–434 K). Single crystals of (I) suitable for X-ray measurements were obtained by recrystallization from cyclohexane at room temperature. $^1\text{H NMR}$ (600 MHz, acetone- d_6): 16.09 (1H, s, OH), 9.68 (1H, s, NH-Ph), 7.06–8.07 (14H, m, Ar). Analysis, calculated for $\text{C}_{21}\text{H}_{16}\text{N}_4\text{OS}$: C 70.76, H 4.52, N 15.72%; found: C 70.81, H 4.60, N 15.68%.

Crystal data

$\text{C}_{21}\text{H}_{16}\text{N}_4\text{OS}$
 $M_r = 372.45$
 Monoclinic, $C2/c$
 $a = 19.089$ (3) Å
 $b = 8.4715$ (11) Å
 $c = 23.161$ (3) Å
 $\beta = 90.315$ (2)°
 $V = 3745.4$ (9) Å³
 $Z = 8$

$D_x = 1.321$ Mg m⁻³
 Mo $K\alpha$ radiation
 Cell parameters from 2238 reflections
 $\theta = 2.6$ – 22.6 °
 $\mu = 0.19$ mm⁻¹
 $T = 293$ (2) K
 Block, colourless
 $0.46 \times 0.31 \times 0.18$ mm

Data collection

Bruker SMART CCD area-detector diffractometer
 φ and ω scans
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)
 $T_{\min} = 0.892$, $T_{\max} = 0.966$
 12 321 measured reflections

4500 independent reflections
 2680 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.023$
 $\theta_{\text{max}} = 28.0$ °
 $h = -24 \rightarrow 25$
 $k = -11 \rightarrow 11$
 $l = -30 \rightarrow 20$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.040$
 $wR(F^2) = 0.121$
 $S = 1.06$
 4500 reflections
 252 parameters
 H atoms treated by a mixture of independent and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0561P)^2 + 0.3455P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.16$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.18$ e Å⁻³
 Extinction correction: none

Table 1

Selected geometric parameters (Å, °).

S1—C1	1.6816 (17)	N2—C2	1.4309 (19)
N1—C1	1.324 (2)	N3—N4	1.2920 (18)
N1—C16	1.428 (2)	N4—C11	1.374 (2)
N2—N3	1.3609 (17)	O1—C3	1.328 (2)
N3—N2—C2—C3	−92.39 (17)	N2—C2—C3—C4	3.6 (2)
N3—N2—C2—C1	84.17 (18)	C1—N1—C16—C17	54.8 (2)

Table 2

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1A \cdots S1	0.88 (2)	2.05 (2)	2.8752 (17)	156 (2)
N1—H1 \cdots N2	0.855 (12)	2.343 (14)	2.7188 (15)	107.0 (11)
N1—H1 \cdots N4 ⁱ	0.855 (12)	2.112 (12)	2.9051 (14)	154

Symmetry codes (i) $-x + 1, y, -z + \frac{1}{2}$.

H atoms on N or O atoms were located in a difference Fourier map and refined freely. All other H atoms were placed in calculated positions, with $C-H = 0.93$ or 0.97 Å, and included in the final cycles of refinement using a riding model, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Data collection: SMART (Bruker, 1998); cell refinement: SAINT (Bruker, 1999); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: SHELXTL (Bruker, 1999); software used to prepare material for publication: SHELXTL.

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References

- Bruker (1998). SMART. Bruker AXS Inc., Madison, Wisconsin, USA.
 Bruker (1999). SAINT and SHELXTL. Bruker AXS Inc., Madison, Wisconsin, USA.
 James, M. N. G. & Watson, K. J. (1966). *J. Chem. Soc. C*, pp. 1361–1371.
 Ji, B. M., Du, C. X., Zhu, Y. & Wang, Y. (2002). *Chin. J. Struct. Chem.* **21**, 252–255.
 Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
 Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.
 Zhang, Y., Sun, X. W., Hui, X. P., Zhang, Z. X., Wang, Q. & Zhang, Q. (2002). *Chin. J. Chem.* **20**, 168–173.